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Research Article

Protective Effect of Alkaloids from *Amaranthus Viridis* Linn. Against Hydrogen Peroxide Induced Oxidative Damage in Human Erythrocytes (RBC)

Abstract

Plants are rich in antioxidants which play an essential role in disease prevention. *Amaranthus viridis* linn. Is a common medicinal plant, spread throughout the world is used in traditional ayurvedic medicine. The present study was aimed to examine the protective effect of partially purified alkaloids (PPA) from *Amaranthus viridis* against hydrogen peroxide induced oxidative damage in human erythrocytes *in vitro* conditions. The experiments are performed in groups (Control, H₂O₂ induced and H₂O₂ + 25, 50 mg conc of alkaloids). The enzymic antioxidants like CAT, SOD, GPx, GST and non enzymic antioxidants like vitamin C, vitamin E and GSH levels and LPO were evaluated to determine the protective effect of alkaloids. The alkaloids prevent the decline of antioxidant status which in turn decreases LPO levels by preventing MDA formation in erythrocytes. The results confirm the protective effect of alkaloids against free radical induced oxidative damage in human erythrocytes.

Abbreviations

CAT: Catalase; SOD: Superoxide Dismutase; GPx: Glutathione Peroxidase; GST: Glutathione-S-Transferase; GSH: Glutathione; LPO: Lipid Peroxidation; MDA: Malondialdehyde; H₂O₂: Hydrogen Peroxide; D.H₂O: Distilled Water; Conc.: Concentrated; ROS: Reactive Oxygen Species; RBC: Red Blood Cells; NADH: Nicotinamide Adenine Dinucleotide; CDNB: 1-Chloro-2,4-Dinitro Benzene; TCA: Trichloro Acetic Acid; EDTA: Ethylene Diamine Tetra Acetate; DTNB: 3,3'-Dithio-Bis(6- Nitrobenzoic Acid); PBS: Phosphate Buffered Saline; Alk: Alkaloid; Hb: Hemoglobin;

Introduction

Free radicals are formed during the normal metabolic process in the biological system. Free radicals are highly reactive due to the presence of unpaired electron(s) [1]. Reactive oxygen species (ROS) are a type of free radicals that are produced during normal cell metabolism, both in animals and plants. The ROS including H₂O₂, superoxide and the hydroxyl radical are some of the metabolic intermediates in human body [2]. Excess of ROS leads to oxidative stress, resulting in oxidative DNA damage which is implicated in the pathogenesis of numerous disorders [3].

A. viridis (Amaranthaceae) is spread throughout the world, growing under a wide range of climatic conditions. They produce useful feed and food products [4]. *A. viridis* L. (Amaranthaceae) has also been used in Indian traditional system. It is also a very good source of vitamins including vitamin A, B6, and C, riboflavin and foliate, minerals like calcium, iron, zinc, magnesium [5]. The *A. viridis* leaf extracts protects the heart by preserving the integrity of membrane and by maintaining the activities of marker enzymes in

the serum and heart of isoproterenol induced cardio toxic rats which may be due to antilipoperoxidative and antioxidant effects [6].

Alkaloids are plant derived pharmacologically active, basic compounds from amino acids that contain one or more heterocyclic nitrogen atoms. Plant derived alkaloids are suggested to be natural antioxidants associated with prevention of various diseases and pathological conditions including malaria, diabetics, cancer, cardiac dysfunction etc. and also used as local anesthetics and pain relievers [7]. Newly extracted alkaloids from medicinal plants are used as therapeutic agents [8]. Many alkaloids are used as anti-arrhythmic, anticholinergic, anti-tumor, vasodilating, antihypertensive, cough medicine, anesthetic, antiprotozoal, antidiabetic, antihyperlipidemic and antioxidant agents [6].

Erythrocytes are mammalian cells that are responsible for transporting oxygen and other substances in blood throughout the body. These erythrocytes are easy to extract and they lack nucleus and hence they can be effectively used for studying the effect of various pharmacologically active substances. They are highly susceptible to oxidative damage as a result of high polyunsaturated fatty acid (PUFA) content of their membranes and the high cellular concentrations of oxygen and hemoglobin (Hb), a potentially powerful promoter of oxidative processes [9].

The present study reveals the protective effect of isolated alkaloids from *A. viridis* leaves against H₂O₂ induced oxidative damage in RBC.

Materials and Methods

Chemicals used

PBS (pH 7), hydrogen peroxide, phosphate buffer, ammonium

hydroxide, potassium dichromate, glacial acetic acid, sodium phosphate buffer, absolute ethanol, chloroform, n-butanol, phenazine methosulphate, nitrobluetetrazolium, ADH, CDNB, reduced glutathione, sodium phosphate buffer, sodium azide, TCA, EDTA, sodium citrate, DTNB, oxalic acid, sulphuric acid, 2,4-dinitro phenyl hydrazine, thiourea, ascorbic acid, disodium hydrogen phosphate, petroleum ether, 2,2-dipyridyl, ferric chloride were purchased from Sigma-Aldrich or Merck. All the chemicals used were of analytical grade.

Preparation of *A. viridis* leaf extract

The fresh plants of *A. viridis* were collected from local market in Erode, Tamil Nadu, India and was authenticated by Dr. R. Jagadesan, Horticulture, Horticultural College and Research Institute for Women, Trichy. The fresh leaves of *A. viridis* were collected and washed with tap water, followed by distilled water and shade dried. The dried leaves were manually ground to a fine powder and stored at 4°C until future use.

Isolation of alkaloids

5g of *A. viridis* powder was weighed in a 250ml conical flask and 50ml of 20% acetic acid in ethanol was added and covered to stand for 4hrs. This was filtered and the extract was concentrated using water bath to one-quarter of the original volume. Concentrated ammonium hydroxide was added drop wise until the precipitation was complete. The whole solution was allowed to settle and the precipitate was collected by filtration and amount of alkaloid yield was weighed.

Purification of alkaloids

The filtered alkaloid was purified using the method proposed by Shi et al., 2002. 5g of silica gel (laboratory grade) were activated in hot air oven at 110°C for 1 hr. The glass wool was fixed at the bottom of the column. The activated silica was added to the column, in small portion with gentle tapping after each addition, in order to ensure the uniform packing. The small quantity of solvent was allowed to remain at the top of the column (about 4 cm). Crude alkaloid residue (0.1g) was loaded to the top the column and eluted with the methanol solvent system. The presence of alkaloids in the extracted fractions were verified by various tests like Dragendorff's test, Wagner's test, Hager's test and Mayer's Test.

Preparation of erythrocytes suspensions [10]

Fresh blood samples from healthy volunteers (10–15ml) were collected and centrifuged at 3000 rpm for 15 minutes, plasma and puffy coats were removed. Red cells were washed with PBS (pH 7.00) for three times and erythrocytes were lysed with ice-cold distilled water.

Experimental design

Erythrocyte suspensions obtained from healthy donor were divided into four groups as Group I- Control [Erythrocyte suspension (750µl), PBS (1000µl) and D.H₂O (250µl)], Group II-H₂O₂ [Erythrocyte suspension (750µl), 10mM H₂O₂ (50µl), PBS (950µl) and D.H₂O (250µl)], Group III- Alk (25 mg) [Erythrocyte suspension (750µl), 10mM H₂O₂ (50µl), alkaloids (500µl) (25mg alkaloids in 500µl PBS) and PBS (950µl)] and Group IV- Alk (50 mg) [Erythrocyte

suspension (750µl), 10mM H₂O₂ (50µl), alkaloids (500µl) (50mg alkaloids in 500µl PBS) and PBS (950µl)]. These experimental groups were incubated at 37°C for 1 hour. Following the incubation, enzymic and non-enzymic antioxidant activity were determined.

Determination of antioxidant status

Measurement of Lipid peroxidation [11]: The lipid peroxidation was measured by the amount of formation of Malondialdehyde (MDA), the end product of lipid peroxidation. The MDA content was determined by the thiobarbituric acid (TBA) reaction. 1.0ml of sample, 1.5 ml of 0.67% TBA and 1.0ml of water was added. A pink colour developed is read at 535nm. A series of standards (2-10 nmoles) was treated in a similar manner along with a reagent blank. The results were expressed for nmoles of MDA formed/mg protein.

Measurement of catalase activity [12]: Catalase causes rapid decomposition of hydrogen peroxide to water. Dichromate in acetic acid was converted to perchloric acid and then to chromic acetate when heated in presence of H₂O₂. The chromic acetate thus produced is measured spectrophotometrically at 610nm. To 0.9ml of phosphate buffer, 0.1ml of sample and 0.4ml of H₂O₂ was added. At 0sec and after 60sec 2.0ml of dichromate-acetic acid mixture was added. The tubes were kept in boiling water bath for 10 min and the colour formed as a result of chromic acetate formation was read at 620nm. Standards in the range of 1.2-6.0 µmol were taken and processed as test and blank containing reagent alone. The activity of catalase was expressed for erythrocytes as µmol of H₂O₂ decomposed/minutes/g of Hb.

Measurement of superoxide scavenging activity [13]: The determination of SOD is based on the inhibition of the formation of NADH-phenazine methosulphate and nitroblue tetrazolium formazon. 0.5ml sample was diluted to 1.0ml with water followed by addition of 2.5ml of ethanol and 1.5ml chloroform (chilled reagents were added). This mixture was shaken for 90 minutes at 4°C and then centrifuged. The assay mixture contained 1.2ml of sodium pyrophosphate buffer, 0.1ml of phenazine methosulphate, 0.3ml of nitroblue tetrazolium and appropriately diluted enzyme preparation in a total volume of 3ml. The reaction was started by the addition of 0.2ml NADH. After incubation at 30°C for 90 min, the reaction was stopped by the addition of 1ml glacial acetic acid. The reaction mixture was stirred vigorously and shaken with 4ml of n-butanol. The mixture was allowed to stand for 10 min, centrifuged and n-butanol layer was separated. The colour density of the chromogen in n-butanol was measured in a Spectrophotometer at 520nm. A series of standard treated in a similar way to determine the enzyme activity. The enzyme concentration required to inhibit the chromogen produced by 50% in one minute under standard conditions was defined as one unit. The specific activity of the enzyme is expressed as unit/minutes/g of Hb.

Measurement of Glutathione-S-transferase activity [14]: Glutathione-S-transferase catalyses the reaction of 1-chloro 2,4 dinitrobenzene (CDNB) with the sulphhydryl group of glutathione. To 1.0ml of buffer, 0.1ml of sample, 1.7ml of water, 0.1ml of CDNB were added and incubated at 37°C for 5 min. After incubation, 0.1ml of reduced glutathione was added. The increase in optical density of the enzyme was measured against blank at 340nm. The enzyme activity was calculated in terms of µmoles of CDNB conjugate formed/minutes/g of Hb.

Estimation of Ascorbic acid [15]: For the estimation of ascorbic acid, 1.0ml sample was taken and made up the volume to 3.0ml with distilled water. Added 1.0ml of dinitro phenyl hydrazine reagent followed by 1 to 2 drops of thiourea into each tube. A blank was done along with standard solution. Mixed the contents and incubated at 37°C for 3 h. After incubation the tubes were kept in the ice bath. Dissolved the orange red osazone crystals formed by adding 7.0ml of 80% sulphuric acid drop wise while the tubes were still in the water bath. The tubes in the ice bath were removed and allowed to stand for 30 min at room temperature and measured the absorbance at 540 nm. The result was expressed as µg/g of Hb.

Estimation of Glutathione [16]: This method was based on the development of yellow colour when 5, 5'Dithio-bis (2-nitrobenzoic acid) (DTNB) was added to compounds containing sulphhydryl groups. 0.5ml of homogenate was pipetted out and precipitated with 2.0ml of 5% TCA.2.0ml of supernatant was taken after centrifugation and 1.0ml of Ellman's reagent and 4.0ml of 0.3 M disodium hydrogen phosphate were added. The yellow colour developed was read in Spectrophotometer at 412nm. A series of standards (20–100µg) was treated in a similar manner along with a blank containing 1.0ml of buffer. The results were expressed in µ mol/g of Hb.

Statistical analysis

The data were expressed as mean ± S.D Statistical differences at p<0.05 between the groups were analyzed by one way ANOVA followed by Dunnett's multiple comparison test using SPSS 15.0 software.

Results and Discussion

Alkaloid yield

The amount of alkaloids obtained from *A. viridis* is found to be 100 mg per 5g of leaves.

Test for alkaloids

The results for qualitative analysis of alkaloids using Dragendorff's test, Wagner's test, Hayer's test, and Mayer's test proved the presence of alkaloids. The test results were presented in **Table 1**.

Antihemolytic activity of alkaloids in H2O2 induced oxidative damage in erythrocytes

Under normal conditions, the continuous production of free radicals is compensated by the powerful action of protective enzymes like superoxide dismutase, catalase and glutathione peroxidase that are believed as major antioxidant enzymes present in the human body that protect against the oxygen toxicity [17]. As a result, a lipid peroxidation (LPO) process occurs. Therefore, antioxidant enzyme activities and lipid peroxidation levels are accepted as important parameters in the evaluation of oxidative stress in aerobic organisms. The current study was carried out with the main aim of evaluating the antioxidant status of the alkaloids that are isolated from the leaves of *Amaranthus viridis*. The partially purified alkaloids from the leaves proved to protect the erythrocytes from the oxidative stress caused by H₂O₂. The results of the present study, i.e., the effect of various antioxidants on H₂O₂ induced oxidative damage on erythrocytes was depicted in Figures: 1-3.

Effects of alkaloids on H₂O₂ induced lipid peroxidation: MDA is formed in high amounts during the lipid peroxidation process and its quantity reveals the extent of cell damage by peroxidation. MDA is a highly reactive, bifunctional molecule that can effectively cross link

Table 1: Qualitative analysis of alkaloids from the sample.

S. No.	TESTS	RESULT
1.	Draggendorf's	++
2.	Wagner's	++
3.	Hayer's	+
4.	Mayer's	++

(++) - Highly present, (+) – Present.

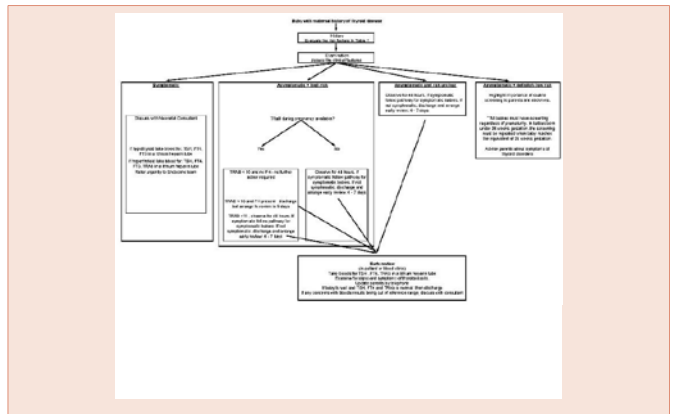


Figure 1: Effect of the plant extract on Lipid peroxidation in erythrocytes.

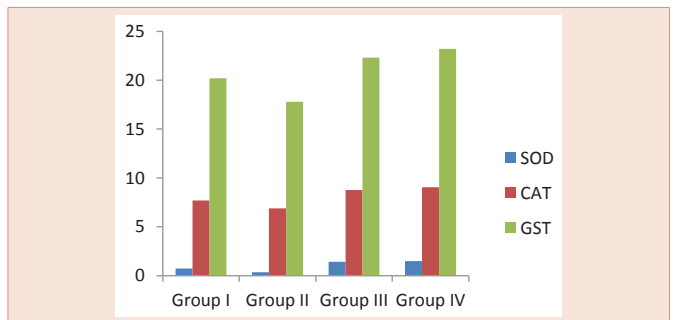


Figure 2: Effect of the plant extract on SOD, CAT and GST activity in erythrocytes.

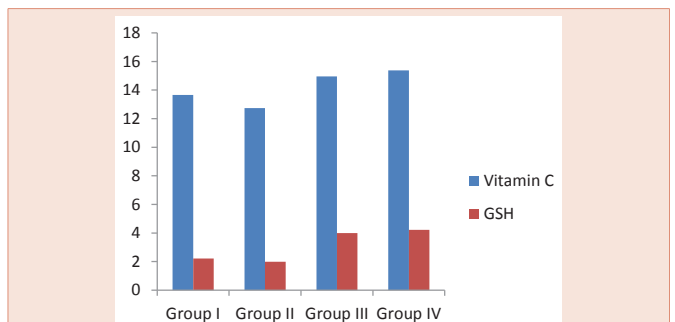


Figure 3: Vitamin C and GSH content in erythrocyte groups.

with the membrane phospholipids and proteins of the erythrocytes, thus impairing the membrane related functions ultimately leading to diminished survival of the erythrocytes [18]. The production of MDA was high in group II where the erythrocyte cells are treated with H₂O₂ alone (Table 2). Treatment with alkaloids prevented H₂O₂ induced MDA production and this inhibition was strongly dose dependent. The values in group IV are reversed and nearer to the control which indicates that alkaloids inhibits the membrane lipid peroxidation triggered by the injurious oxygen radicals generated from hydrogen peroxide (Figure 1).

Effects of alkaloids on CAT, SOD and GST activity: Enzymic antioxidants, such as catalase, superoxide dismutase and glutathione S-transferase, play important roles in the scavenging of reactive oxygen species, such as superoxide radical, hydrogen peroxide, lipid hydroperoxides, and so forth. An aerobic organism can cope with the metabolic production of ROS under normal conditions via its antioxidant defense system; however, overproduction of ROS causes a dangerous process called “oxidative stress” [19]. The antioxidant effect of CAT, SOD and GST on erythrocyte suspensions were presented in Table 3. The effect of all these three antioxidants, viz., CAT, SOD and GST was lowered in group II where H₂O₂ induced oxidative damage was high. The group IV samples with high concentration of the purified alkaloid showed more protective effect against oxidative damage caused by H₂O₂. The activity of the enzymic antioxidants in presence of the alkaloids was high and it is dose dependent, higher the alkaloid then there is higher free radical scavenging activity (Figure 2).

Effects of alkaloids on Ascorbic acid and Glutathione: Vitamin C and Glutathione are some of the non-enzymic antioxidants of animals which help those organisms to combat cellular damage due to oxidative stress. Vitamin C may accumulate at high concentration in photosynthetic tissues, in which it is intimately in the regulation of photosynthesis and protect the chloroplasts against damage caused by ROS such as O₂, H₂O₂, hydroxyl radicals (OH) and singlet oxygen [20]. The effect of non-enzymic antioxidants on the erythrocyte suspension was presented in Table 4. Vitamin C is proved to have higher free radical scavenging activity in presence of the partially purified alkaloid from *A. viridis*. The protective effect of those non-enzymic antioxidants was preserved even in presence of H₂O₂ and their activity was nearly normal to that of the control sample (Figure

3). Glutathione is also a non-enzymic antioxidant which also possesses higher free radical scavenging activity and its antioxidant effect was found to be preserved at higher rates in presence of increased amount of the alkaloid from the leaf extract.

Conclusion

The antihaemolytic activity of the alkaloids from *A. viridis* was confirmed from this study. The results showed alkaloids present in the plant have good antihaemolytic activity. As a conclusion from the present study the alkaloids from *A. viridis* were found to have more protective effects on the antioxidant systems against H₂O₂ induced oxidative damage on erythrocytes. These alkaloids can effectively be used as a natural antioxidant for the treatment and prevention of lipid peroxidation related diseases.

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Table 2: Lipid peroxidation.

S. No.	GROUPS	LPO
1.	I	0.90± 0.10**
2.	II	1.2 ±0.16 **
3.	III	1.16 ±0.14**
4.	IV	0.64±0.13**

Table 3: Enzymic antioxidant activity in erythrocytes.

S.NO	Groups	SOD (50% inhibition/ min/g Hb)	CAT (μ moles/ min/g Hb)	GST (μ moles/ min/g Hb)
1.	I	0.72±0.22**	7.69±0.45**	20.19±0.82***
2.	II	0.35±0.10*	6.88±0.16**	17.8±0.53*
3.	III	1.42±0.25*	8.75±0.54**	22.3±1.02**
4.	IV	1.48±0.12**	9.04±0.75**	23.21±1.05*



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